

PAPER



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The phytase RipBL1 enables the assignment of a specific inositol phosphate isomer as a structural component of human kidney stones†

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Inositol phosphates (InsPs) are ubiquitous in all eukaryotes. However, since there are 63 possible different phosphate ester isomers, the analysis of InsPs is challenging. In particular, InsP₁, InsP₂, and InsP₃ already amass 41 different isomers, of which some occur as enantiomers. Profiling of these “lower” inositol phosphates in mammalian tissues requires powerful analytical methods and reference compounds. Here, we report an analysis of InsP₂ and InsP₃ with capillary electrophoresis coupled to electrospray ionization mass spectrometry (CE-ESI-MS). Using this method, the bacterial effector RipBL1 was analyzed and found to degrade InsP₆ to Ins(1,2,3)P₃, an understudied InsP₃ isomer. This new reference molecule then aided us in the assignment of the isomeric identity of an InsP₃ while profiling human samples: in urine and kidney stones, we describe for the first time the presence of defined and abundant InsP₃ isomers, namely Ins(1,2,3)P₃, Ins(1,2,6)P₃ and/or Ins(2,3,4)P₃.

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Introduction

myo-Inositol phosphates (InsPs) are molecules with various numbers of phosphate groups modified on one to six of the OH groups of *myo*-inositol (hereafter Ins), which are present in all eukaryotes. By sequential phosphorylation, 63 different InsPs can in principle be generated, of which over one third are thought to be relevant to mammalian metabolism as signalling molecules.^{1,2}

Analysis of these InsPs is important to better characterize InsP identity and abundance in mammalian metabolism, which might finally lead to a deciphering of the alleged “inositol phosphate

code”.^{1,3,4} A variety of analytical methods have been reported, including, for example, strong anion exchange high performance liquid chromatography (SAX-HPLC), high performance liquid chromatography coupled to mass spectrometry (HPLC-MS), and capillary electrophoresis coupled to electrospray ionization mass spectrometry (CE-ESI-MS).^{5–13} Most of these methods are sensitive and efficient for the analysis of highly phosphorylated InsPs (*e.g.* InsP₄, InsP₅, InsP₆).^{5–11} Moreover, CE-MS methods are powerful tools for profiling of inositol pyrophosphates (PP-InsPs), such as InsP₇ and InsP₈ that carry one or several diphosphate groups.^{6,14–16}

Analyzing identity and abundance of InsP₁, InsP₂ and InsP₃ isomers in biological samples remains a significant challenge for several reasons. First, due to the non-specificity of the current extraction and separation methods, it is difficult to identify InsP₁ and InsP₂ isomers from the isobaric and more abundant sugar mono- and diphosphates. Second, the number of possible InsP₁ to InsP₃ isomers amasses to 41 and out of these alone 20 isomers of InsP₃ exist, including 8 enantiomeric pairs. Enantiomer separation on a chiral stationary phase remains an unsolved issue for these molecules. Additionally, it is generally assumed that the *myo*-configuration is the relevant one, but also other inositol configurations occur in biology.¹⁷ Moreover, the high negative charge density and the absence of a chromophore generate significant issues regarding sensitive detection. The most commonly used analytical methods for InsP₂ and InsP₃ separation are SAX-HPLC and

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HPLC-MS.^{9,18,19} The latter can profit from a heavy isotope labelling methylation strategy to enable ESI⁺ measurements.⁹ Ins(1,4,5)P₃, the Ca²⁺ release factor,²⁰ is the most characterized InsP₃ and represents the textbook example of a second messenger. Other InsP₃ isomers have been described in biology, such as Ins(1,3,4)P₃ and Ins(1,2,3)P₃.¹⁸ InsP₍₁₋₅₎ with a phosphate group in the 2-position are usually not described in mammalian metabolite overviews,^{2,21-24} although such esters of InsP₂ and InsP₃ have been found since 1995.^{18,25,26} Recently, using ¹³C enrichment and 2D-NMR analysis, it was discovered that Ins(2,3)P₂ and Ins(2)P are major metabolites in immortalized mammalian cell lines, calling into question the general notion that InsPs with a phosphate in the axial 2-position are biologically irrelevant.²⁷ A recent finding that 2-PP-InsP₅ carrying a pyrophosphate in the 2-position is a biologically relevant species, further underscores the need to reassess inositol phosphate structures.¹⁶ Recently developed new analytical technologies are now able to reveal several 'discarded' isomers, whose biological importance must now be evaluated.

Our approach to analyze InsP₄₋₆ and PP-InsPs by CE-ESI-MS is extended herein to study InsP₂₋₃. We were interested to profile human samples and since the involvement of PP-InsPs in systemic human phosphate homeostasis has become clearer,²⁸⁻³¹ we decided to analyze urine samples where excess phosphate is excreted. Surprisingly, we detected several InsP₃ isomers in urine. This led us to postulate that these InsP₃ might be structural components of kidney stones. If this would be the case, we could potentially use urine InsP presence as a biomarker for impending stone formation. To achieve our objectives, we reveal that RipBL1, a bacterial phytase effector protein,³² selectively dephosphorylates InsP₆ to Ins(1,2,3)P₃. Our CE-ESI-MS method combined with [¹³C₆] InsP₃ produced from [¹³C₆] InsP₆^{33,34} by RipBL1 now enables us to assign InsP₃ in human kidney stone and urine and demonstrates in patient samples (healthy vs. kidney stone formers) that the most abundant isomers host a phosphate in the 2-position.

Results

InsP₂ and InsP₃ isomers are well separated by CE-ESI-MS

The CE-ESI-MS analytical method has been originally developed to investigate inositol pyrophosphate metabolism and is a particularly effective separation platform for isomers of InsP₄₋₆ and also for PP-InsPs.⁶ CE-ESI-MS was not further developed to analyze InsP₁₋₃ in biological samples, despite the potential for the determination of spiked InsP₃ in plasma by capillary zone electrophoresis-mass spectrometry.¹² There are 15 isomers of InsP₂ that include 6 enantiomeric pairs, and 20 InsP₃ isomers that include 8 enantiomeric pairs. We disclose a CE-ESI-MS method for analyzing six commercially available isomers of InsP₂ and seven commercially available isomers of InsP₃ by baseline separation, other than enantiomers, which one cannot discriminate by using an achiral bare fused silica capillary.

The set of commercial InsP₂ and InsP₃ are shown in Fig. 1A: Ins(4,5)P₂, Ins(1,2)P₂, Ins(1,5)P₂, Ins(2,4)P₂, Ins(1,3)P₂, Ins(1,4)P₂,

Ins(3,4,5)P₃, Ins(1,2,6)P₃, Ins(1,3,4)P₃, Ins(1,4,5)P₃, Ins(2,4,5)P₃, Ins(1,3,5)P₃, Ins(2,3,5)P₃. A bare fused silica capillary with a length of 100 cm was implemented for separations by applying 30 kV across the capillary (Agilent 7100 CE). Detection was achieved with an ESI-QQQ-MS (Agilent 6495C Triple Quadrupole with Agilent Jet Stream electrospray ionization source) in the negative ionization mode connected with an Agilent CE-ESI-MS interface. Initially, 35 mM ammonium acetate titrated with ammonium hydroxide to pH 9.75 was used as background electrolyte (BGE). With this BGE, separation of six InsP₂ isomers is achieved, but Ins(1,5)P₂, Ins(2,4)P₂, and Ins(1,3)P₂ were not baseline separated (Fig. 1B). Also, the separation of eight InsP₃ isomers was achieved with the exception of Ins(1,3,4)P₃, Ins(1,4,5)P₃, and Ins(1,4,6)P₃ (Fig. 1C). After pH optimization and BGE screening, a near baseline separation of Ins(1,5)P₂ and Ins(2,4)P₂ was achieved with 50 mM ethylamine titrated with formic acid to pH 10.0 (Fig. 1D). Separation of previously coeluting Ins(1,3,4)P₃ and InsP(1,4,5)P₃ was also resolved using this BGE (Fig. 1D) and a baseline separation of InsP(1,3,5)P₃ and Ins(2,3,5)P₃ was achieved as well (Fig. 1D). Ins(1,4,6)P₃ still coeluted with Ins(1,4,5)P₃.

Method validation was performed with InsP standards, including linearity, limit of detection (LOD) and limit of quantification (LOQ). The external calibration curves of InsP₂ and InsP₃ were constructed at eight concentration levels by regression of concentrations against the analyte peak area (Fig. S1, ESI⁺). The calibration curves were linear and had a coefficient of determination > 0.997 over the investigated range of 0.1–10 µg mL⁻¹ for Ins(1,2)P₂ and 0.4–40 µg mL⁻¹ InsP(1,2,6)P₃. With 20 nL sample injection, the LODs are 0.025 µg mL⁻¹ for InsP₂ (*i.e.* 1.3 fmol) and 0.020 µg mL⁻¹ for InsP₃ (*i.e.* 0.8 fmol), the LOQs are 0.05 µg mL⁻¹ for InsP₂ (*i.e.* 2.6 fmol) and 0.10 µg mL⁻¹ for InsP₃ (*i.e.* 4 fmol; Fig. S1, ESI⁺). The limit of detection of this method is significantly lower compared to other chromatography-related methods summarized in ref. 35.

RipBL1 degrades InsP₆ to Ins(1,2,3)P₃

RipBL1 is an effector protein of the phytopathogenic Gram-negative bacteria *Ralstonia solanacearum*, that shares homology with the bacterial effector XopH from *Xanthomonas campestris*. Previous work demonstrated that XopH displays an unusual phytase activity that dephosphorylates InsP₆ stereoselectively at the C1 position only, resulting in the generation of Ins(2,3,4,5,6)P₅.³² We tested InsP₆ dephosphorylation by phytases, such as RipBL1 *in vitro*, to see if we could generate more high-quality non-commercial standards. Importantly, after 45 min 100% of InsP₆ was degraded by RipBL1, and one defined InsP₃ was detected as constituting about 95% of the digestion product. There was very little dephosphorylation to InsP₄ and InsP₂ (Fig. S2B, ESI⁺). CE-qTOF-MS analyses of the enzymatic reaction product confirmed a main molecular ion peak at *m/z* 418.9551 corresponding to InsP₃ (Fig. S2A, ESI⁺). To structurally assign the InsP₃ product, we generated a [¹³C₆] InsP₃ by [¹³C₆] InsP₆³³ digestion with RipBL1, and then analyzed its identity through spiking-in of all commercially available InsP₃ standards. As shown in Fig. 2A, the [¹³C₆] InsP₃ was

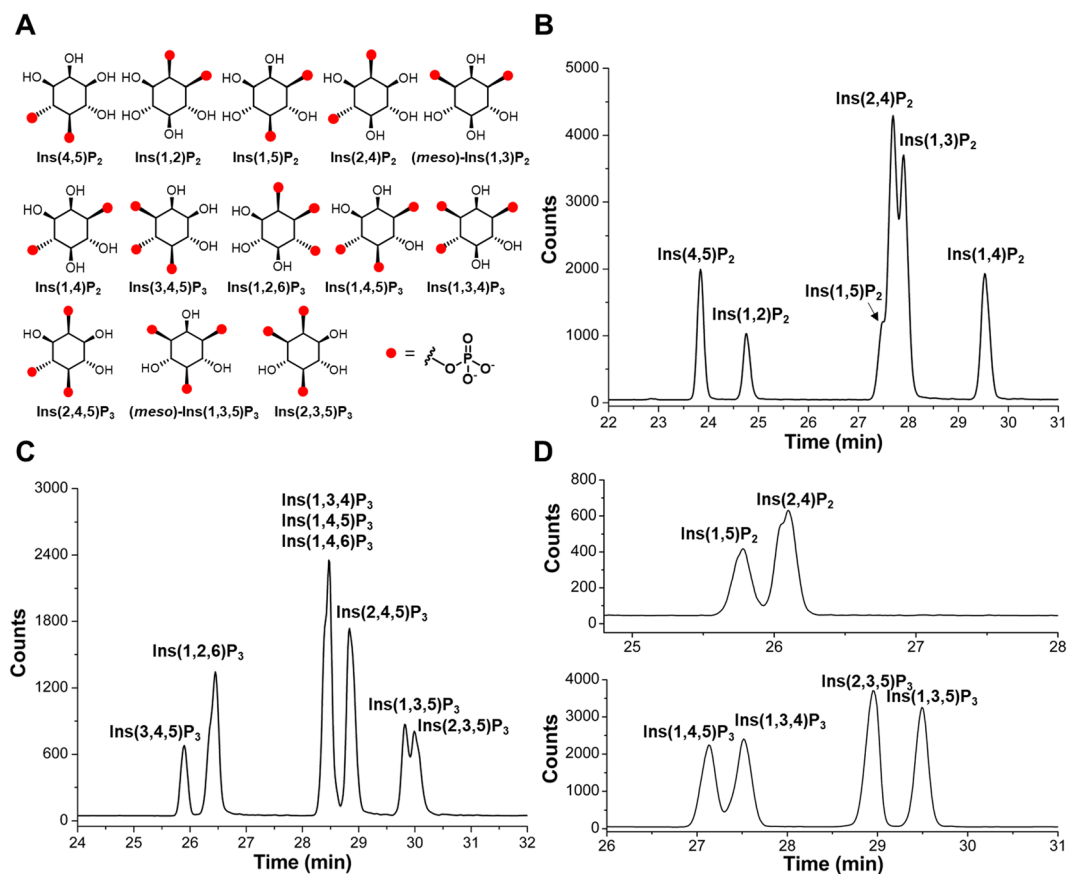


Fig. 1 Separation of InsPs by CE-ESI-MS. (A) Structures of commercial InsP₂ and InsP₃ isomers. Achiral isomers are labeled as “meso”. (B and C) Separation of InsP₂ and InsP₃ standards by CE-ESI-MS, BGE: 35 mM ammonium acetate titrated with ammonium hydroxide to pH 9.75. (D) Separation of InsP₂ and InsP₃ standards by CE-ESI-MS, BGE: 50 mM ethylamine titrated with formic acid to pH 10.0.

baseline separated from all commercially available InsP₃ standards, with the exception of Ins(3,4,5)P₃ but also for this isomer, there was no perfect comigration. Previous work suggests that pyrohydrolysis can partially degrade “higher” InsPs and generate “lower” InsPs through dephosphorylation, and the pyrohydrolysis does not cause phosphate migration.^{19,36} Analysis of pyrohydrolysis products of Ins(3,4,5)P₃ and the [¹³C₆] InsP₃ solution after heating to 100 °C for 2.5 h confirmed that [¹³C₆] InsP₃ is different from Ins(3,4,5)P₃, as the labeled vs. non-labeled InsP₂ products were different (Fig. 2B).

As Ins(4,5,6)P₃, Ins(2,3,6)P₃ (or its enantiomer Ins(1,2,4)P₃), as well as the meso-compounds Ins(1,2,3)P₃, and Ins(2,4,6)P₃ were not readily available, the further assignment was performed differently. We compared the migration time of [¹³C₆] InsP₃ with the pyrohydrolysis products of Ins(2,3,5,6)P₄ and Ins(1,4,5,6)P₄ solution, respectively. The [¹³C₆] InsP₃ product of RipBL1 does not comigrate with any InsP₃ isomers found as pyrohydrolysis products of Ins(2,3,5,6)P₄ or Ins(1,4,5,6)P₄ (Fig. S3, ESI[†]). These results provide evidence that the identity of the [¹³C₆] InsP₃ product is either Ins(1,2,3)P₃ or Ins(2,4,6)P₃, both of which are symmetric meso-compounds.

Pyrohydrolysis products of Ins(1,2,3)P₃ and Ins(2,4,6)P₃ are expected to be different. For Ins(1,2,3)P₃, three InsP₂, namely Ins(1,2)P₂, Ins(2,3)P₂ and Ins(1,3)P₂, should be the pyrohydrolysis

products. On the other hand, Ins(2,4)P₂, Ins(2,6)P₂, and Ins(4,6)P₂ would be the expected pyrohydrolysis products of Ins(2,4,6)P₃. By comparing the electrophoretic mobility with the commercially available InsP₂ standards, two InsP₂ peaks were identified as Ins(1,2)P₂ and its enantiomer Ins(2,3)P₂ (major product) as well as Ins(1,3)P₂ (minor product) in the pyrohydrolysis solution of [¹³C₆] InsP₃ (Fig. 2C). Consequently, [¹³C₆] InsP₃ generated by RipBL1 treatment of [¹³C₆] InsP₆ represents Ins(1,2,3)P₃.

Dephosphorylation of InsP₆ by different types of phytases to produce Ins(1,2,3)P₃, either as the final product or as intermediates, has been investigated.^{37–41} Ins(1,2,3)P₃ as an intermediate in barley aleurone tissue was described as well.⁴² That InsP₆ metabolism by phytase in plants and fungi might pass through Ins(1,2,3)P₃ was discussed.¹⁸ Ins(1,2,3)P₃ as the final product of InsP₆ hydrolysis, as shown for RipBL1, is however unusual, only alkaline phytase from Lily Pollen is a candidate enzyme but the Ins(1,2,3)P₃ purity was not described.³⁸ Thus, to the best of our knowledge, RipBL1 is the first bacterial effector phytase that with high selectivity and yield generates Ins(1,2,3)P₃.

Ins(1,2,3)P₃, Ins(1,2,6)P₃ and/or its enantiomer Ins(2,3,4)P₃ exist in human kidney stones and urine

Kidney stones are solid objects in the kidney or bladder that can cause different disease symptoms and consist of various low

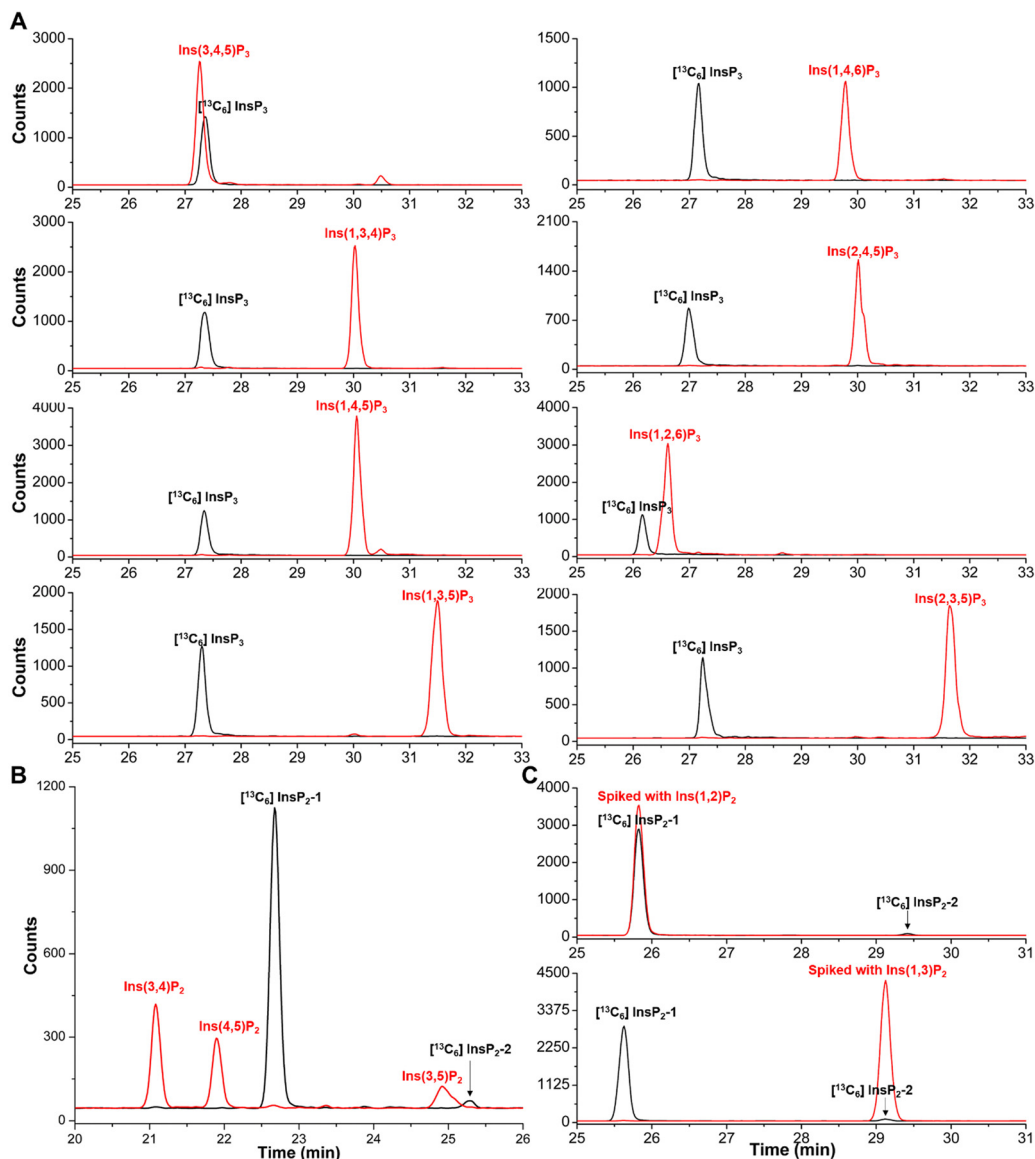


Fig. 2 Identification of the $InsP_3$ (black line) dephosphorylation product of $[^{13}C_6]$ $InsP_6$ by the RipBL1 enzyme. (A) CE-ESI-MS analysis of $InsP_3$ individually spiked (red line) with $Ins(3,4,5)P_3$ standard, $Ins(1,4,6)P_3$ standard, $Ins(1,3,4)P_3$ standard, $Ins(2,4,5)P_3$ standard, $Ins(1,4,5)P_3$ standard, $Ins(1,2,6)P_3$ standard, $Ins(1,3,5)P_3$ standard or $Ins(2,3,5)P_3$ standard, as indicated. (B) Analysis of $InsP_2$ generated from $Ins(3,4,5)P_3$ and the $[^{13}C_6]$ $InsP_3$ isomer by heating to 100 °C for 2.5 h. Extracted ion electropherograms of $[^{13}C_6]$ -labelled $InsP_2$ (black lines) generated by $[^{13}C_6]$ $InsP_3$ and $InsP_2$ (red trace) generated by $Ins(3,4,5)P_3$. (C) $[^{13}C_6]$ $InsP_2$ (black line) generated from $[^{13}C_6]$ $InsP_3$ after heating to 100 °C for 2.5 h spiked with $Ins(1,2)P_2$ standard or $Ins(1,3)P_2$ standard as indicated (red line). Note that the $[^{13}C_6]$ $InsP_2-1$ isomer has the same migration time as the $Ins(1,2)P_2$ standard and that the $[^{13}C_6]$ $InsP_2-2$ isomer has the same migration time as the $Ins(1,3)P_2$ standard.

molecular weight compounds as well as proteins.⁴³ Further, kidney stones can be categorized into calcium containing stones and non-calcium containing stones. Calcium containing stones are the most common forms of kidney stones, including calcium oxalate monohydrate (COM) or dihydrate (COD) as well as calcium phosphate and mixtures of these.⁴³ The formation of kidney stones is mainly driven by urinary supersaturation and crystallization. These processes are environment dependent, influenced by urine pH, concentration of specific substances and effective molecules (promoters, and inhibitors of kidney

stone formation).⁴⁴ Studies have shown that for example urinary $InsP_6$ can inhibit crystallization during the process of kidney stone formation in a model system.^{45–48} Recently, $InsP_6$ analogues with PEG modifications were reported to completely inhibit such crystallization processes in the nanomolar range.⁴⁹ Additionally, studies with SNF472 (a hexasodium salt of $InsP_6$) as an inhibitor of vascular calcification in a phase 2 clinical trial are ongoing,⁵⁰ highlighting a strong relationship between $InsP_6$ and calcification. Whether some other inositol phosphates might actually promote crystallization and whether these

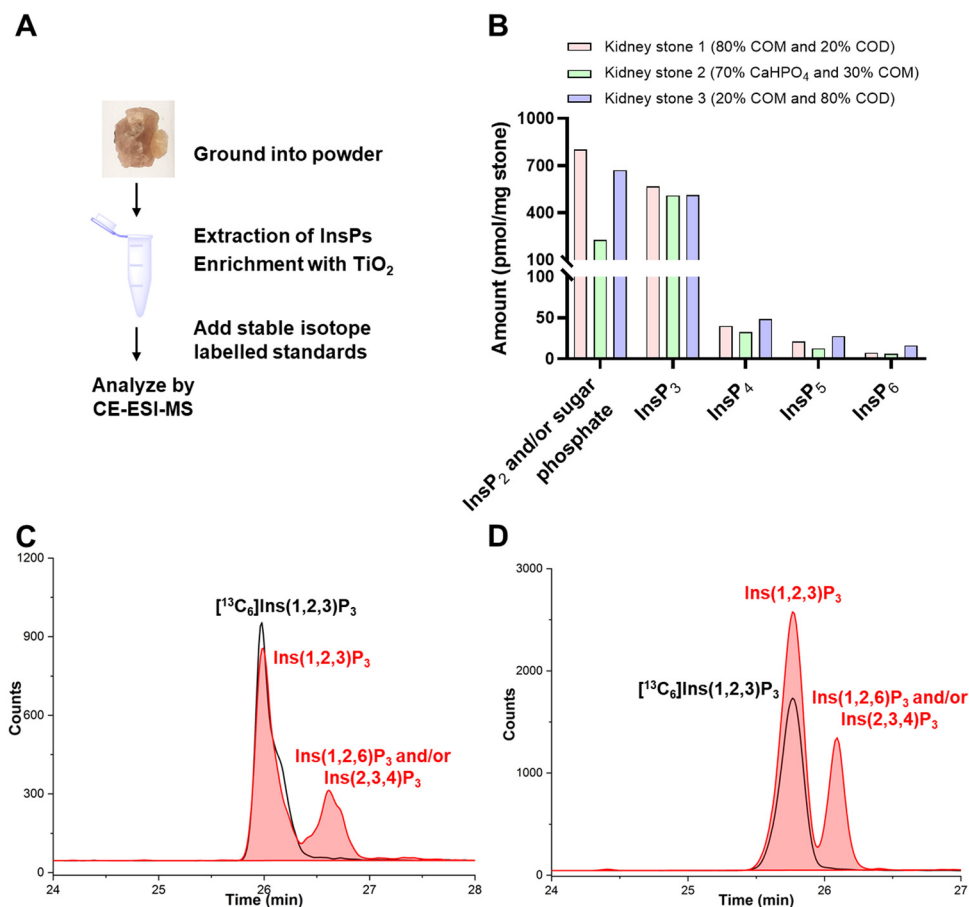


Fig. 3 Profiling of InsPs in human kidney stones. (A) Extraction and analysis workflow of kidney stones for CE-ESI-MS. (B) InsPs distribution in kidney stone 1 (contains 80% COM and 20% COD), kidney stone 2 (contains 70% CaHPO₄ and 30% COM) and kidney stone 3 (contains 20% COM and 80% COD) from three different patients. (C) Extracted ion electropherograms of [¹³C₆] Ins(1,2,3)P₃ (black line) and InsP₃ in kidney stone 1 (red area). (D) Extracted ion electropherograms of [¹³C₆] Ins(1,2,3)P₃ (black line) and InsP₃ in kidney stone 2 (red area).

substances are then incorporated into the kidney stone has not been studied. We therefore asked the question if InsP₆ or other InsPs could be structural components of kidney stones.

To investigate the presence and profiles of InsPs in kidney stones, different ground kidney stones of several compositions (calcium oxalate * xH₂O (x = 1 COM, x = 2 COD) and calcium hydrogen phosphate (CaHPO₄)), determined by IR spectroscopy, were extracted with perchloric acid according to the reported TiO₂ purification method⁵¹ and then analyzed by CE-ESI-MS (Fig. 3A).⁶ We identify several InsPs as part of kidney stones providing the first such profiles (for a representative example see Fig. S4, ESI†). The high-resolution masses of the InsPs were confirmed by a CE-qTOF analysis (Fig. S5, ESI†). Some of the InsP isomers were identified with internal [¹³C₆] labelled reference compounds. Fig. 3B summarizes the observed levels of different InsPs in calcium oxalate stones (kidney stone 1 contains 80% COM and 20% COD, kidney stone 3 contains 20% COM and 80% COD) and a calcium phosphate stone (kidney stone 2 contains 70% CaHPO₄ and 30% COM), which roughly show the same trend in abundance of the analytes.

While inositol pyrophosphates were not detectable in these three types of calcium containing kidney stones, we noticed a decrease in InsP abundance proportional to the number of

phosphate groups. InsP₆ was the least abundant InsP followed by InsP₅ and InsP₄ isomers. One of the InsP₅ isomers was assigned as 2-OH InsP₅ (Fig. S4, ESI†) by spiking with an internal [¹³C₆] reference. The most abundant InsP detected in kidney stone were InsP₂₋₃ species and or isobaric sugar bisphosphates.

Full recovery for InsP₆₋₇ and good recovery for 2-OH InsP₅ from mammalian cell extracts were reported previously.⁶ Here our analysis shows good recovery for InsP₆ (84%), 2-OH InsP₅ (70%) and InsP₃ (71%) from kidney stone extracts (kidney stone 2 is used as a representative example) by spiking with the [¹³C₆] reference before extraction (pre-spiking) and after extraction but before measurement (post-spiking) (Fig. S6, ESI†). Since kidney stones are rich in Ca²⁺, which could affect InsP recovery, we reassessed TiO₂ mediated InsP retrieval by adding 15 mM ethylenediaminetetraacetic acid (EDTA) during extraction. However, the recovery of InsP₆ and InsP₃ did not critically rely on presence or absence of additional EDTA.

Three peaks belonging to InsP₂ and/or isobaric sugar bisphosphates were separated in kidney stone (Fig. S7, ESI†). Spiking indicated that none of these three peaks represents glucose-1,6-bisphosphate. The most intense peak has an

identical migration time with [$^{13}\text{C}_6$] Ins(1/3,2)P₂ generated by pyrohydrolysis from [$^{13}\text{C}_6$] Ins(1,2,3)P₃ (Fig. S7, ESI[†]), indicating Ins(1,2)P₂ and/or Ins(2,3)P₂ are present in kidney stones. This is in line with recent findings that Ins(2,3)P₂ is present in the μM range in immortalized mammalian cells.²⁷

InsP₃ was also comparably abundant in the calcium containing stones. Two peaks of InsP₃ were recorded and identified as Ins(1,2,3)P₃, representing the most intense peak, and Ins(1,2,6)P₃ and/or its enantiomer Ins(2,3,4)P₃. Identification was achieved by spiking with [$^{13}\text{C}_6$] Ins(1,2,3)P₃ (Fig. 3C and D) and Ins(1,2,6)P₃ (Fig. S8, ESI[†]). Ins(1,2,3)P₃ was reported in mammalian cell models in a concentration range of 1–10 μM .¹⁸ A possible role for Ins(1,2,3)P₃ as an intracellular iron chelator in the process of iron transport has been considered.^{52,53} Ins(1,2,6)P₃ and/or the enantiomer Ins(2,3,4)P₃ was proposed in mammalian B-cells in 1992,⁵⁴ however, as discussed above, since then very few studies have been conducted to characterize their identity and functional roles and these isomers are missing in discussions in recent literature reviews.^{21,22}

The formation of kidney stone is a result of urinary supersaturation and crystallization. To study potential correlations of the profiles of InsPs in kidney stones and urine, we additionally profiled InsPs in 0.4 mL urine samples both from patients who have kidney stones (9 urine samples from different donors) and matched healthy people (10 urine samples from different donors, Table S1, ESI[†]). The accurate masses of InsPs identified in urine samples were confirmed by a CE-qTOF analysis (for representative example see Fig. S9 and S10, ESI[†]). Since we cannot yet distinguish sugar mono and bisphosphates from InsP₁ and InsP₂ because of limitations of our current method, we must assume that m/z 259.0229 and m/z 338.9889 correspond to InsP₁ and/or sugar phosphates and InsP₂ and/or sugar bisphosphates, respectively. CE-QQQ was then used to profile InsP levels also in urine samples. The CE-QQQ results indicated that InsP₂ and/or sugar bisphosphates are the most abundant species, followed by InsP₁ and/or sugar phosphates. Interestingly, InsP₃

was in the same concentration range as the sugar phosphates/InsP₁ group of analytes and InsP₄ was also present in the samples, but much less concentrated (*ca.* 7–8 fold) (Fig. 4A).

InsP₆ was detectable in five of the urine samples from ten healthy people and only one of the urine samples from nine kidney stone patients. In all cases, the concentration of InsP₆ was lower than the limit of quantification (LOQ). 2-OH InsP₅ was also in between the LOD and LOQ in seven of the urine samples from ten healthy people and only three of the urine samples from nine kidney stone patients. According to the signal-to-noise of spiked 1 μM [$^{13}\text{C}_6$] InsP₆ and [$^{13}\text{C}_6$] 2-OH InsP₅, we determined limits of detection (LOD) of InsP₆ and 2-OH InsP₅ in the urine samples analyzed were approx. 7.5 nM and 1.5 nM InsP₅, respectively. Limits of quantitation (LOQ) of InsP₆ and 2-OH InsP₅ were approx. 22.5 nM and 4.5 nM, respectively. Therefore, less than 7.5 nM of InsP₆ exist in most of these urine samples, which is in accordance with the reported detection using InsPs specific assays.^{51,55,56} In half of these samples, approx. 1.5 nM to 4.5 nM 2-OH InsP₅ exist. InsP₄ were also present in urine samples. InsP₅ and InsP₄ received little attention so far in human urine, and our results provide an initial overview of these species. A study of InsPs in rat urine after giving InsP₆ dietary treatment indicated that InsP₆ as well InsP₂, InsP₃, InsP₄ and InsP₅ is excreted in the urine.⁵⁵ The source of those isomers remains obscure.

Similar to the results obtained for kidney stones, InsP₂ and/or sugar bisphosphates are most abundant. The most intense species comigrates with [$^{13}\text{C}_6$] Ins(1/3,2)P₂ generated by pyrohydrolysis of [$^{13}\text{C}_6$] Ins(1,2,3)P₃ (Fig. S9, ESI[†]). Also in agreement with the results found in kidney stones, two peaks corresponding to the mass of InsP₃ in urine samples can be identified as Ins(1,2,3)P₃ and Ins(1,2,6)P₃ and/or the enantiomer Ins(2,3,4)P₃ (Fig. 4B, C and Fig. S11, ESI[†]). According to our current data set, there is no significant difference of InsPs and/or sugar phosphates in urine from patients and healthy people regarding levels of InsP₁ to InsP₄. For InsP_{5–6} the picture is less

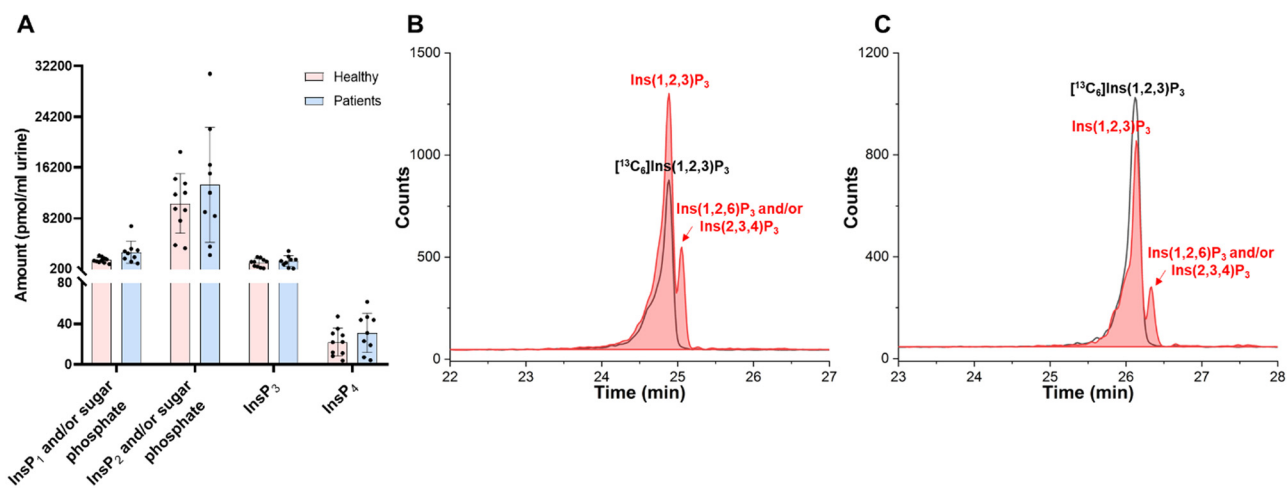


Fig. 4 InsPs in urine samples from patients with kidney stones vs. healthy people. (A) InsPs distribution in urine samples: ten samples from healthy individuals, and nine samples from patients with kidney stones. (B) Extracted ion electropherograms of [$^{13}\text{C}_6$] Ins(1,2,3)P₃ (black line) and InsP₃ in urine from patients (red trace). (C) Extracted ion electropherograms of [$^{13}\text{C}_6$] Ins(1,2,3)P₃ (black line) and InsP₃ in urine from healthy individuals (red trace).

clear as the measured concentrations were in between our current LOD and LOQ for most samples.

Conclusions

We have developed a CE-MS method to separate “lower” InsPs, particularly focusing on InsP₂ and InsP₃. Using this method combined with a pyrohydrolysis strategy, we were able to delineate the identity of a major human InsP₃, *i.e.* Ins(1,2,3)P₃ in urine and kidney stones. To achieve this goal, a phytase (RipBL1) was used that selectively degrades InsP₆ to Ins(1,2,3)P₃ and that also served for production of a ¹³C labelled Ins(1,2,3)P₃ internal reference. Our profiling of InsPs in human kidney stone and urine also unveiled Ins(1,2,6)P₃ and/or the enantiomer Ins(2,3,4)P₃ besides the major Ins(1,2,3)P₃. Importantly, the assignments of InsP₃ are based on an accurate mass determination and comigration with standards. As a next step, further studies must delineate the stereoisomeric identity of the newly identified InsP₃, since chiral selectors have already been developed for certain InsP enantiomers for assignments by ³¹P NMR spectroscopy.³² Additionally, methods to distinguish InsP₁ and InsP₂ from sugar (bis)phosphates will have to be developed. Derivatization of sugar phosphates to separate them from InsPs could be considered.⁵⁷ Further improvements in LOQs will be helpful to establish, if potentially InsP₅ or InsP₆ concentrations in urine can be used as biomarkers for kidney stone formation.

Ins(1,2,3)P₃ was described in mammalian cells more than 25 years ago^{18,26} and other inositol phosphates with a phosphate ester in the 2-position clearly exist. However, the interest in these isomers in human biology has faded over time, and recent literature is not reporting on them anymore so their roles remain unresolved. Our study in combination with work from the Fiedler group²⁷ demonstrates that Ins(2)Ps are abundant cellular, kidney stone and body fluid components. These discoveries warrant reassessment of the classically discussed InsP metabolism. New attention must be given to noncanonical InsP species to fully and properly appreciate the inositol phosphate code.

Author contributions

G. L., A. S., and H. J. J. designed research and wrote the paper. G. L., D. Q. and H. J. J. analyzed data, E. R., R. S., G. S., A. S., T. L. cloned and analyzed RipBL1, D. C., O. B., C. A. W., J. P. J., T. K. provided materials and analyzed data. All authors discussed the paper and provided input for the final version. Further Swiss Kidney Stone Cohort investigators are listed in the acknowledgments.

Conflicts of interest

The authors declare no conflicts of interest.

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